<http://www.pbslearningmedia.org/asset/lsps07_int_ionicbonding/> Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Ionic Bonding Notes:

Like charges \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Draw two examples \_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_

Opposite charges \_\_\_\_\_\_\_\_\_\_\_\_\_ Draw one example \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

The positive ions ( \_\_\_\_\_\_\_\_\_) tend to lose \_\_\_\_\_\_\_\_\_\_\_ and are the \_\_\_\_\_\_\_\_\_\_ which are located on the \_\_\_\_\_\_\_ \_\_\_\_\_\_\_ of the periodic table.

The negative ions ( \_\_\_\_\_\_\_\_\_) tend to gain\_\_\_\_\_\_\_\_\_\_\_ and are the \_\_\_\_\_\_\_\_\_\_ which are located on the \_\_\_\_\_\_\_ \_\_\_\_\_\_\_ of the periodic table.

Which group is the exception to giving away or gaining electrons? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Explain why: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

What happens when a sodium atom and chlorine atoms combine as ions?

Sodium \_\_\_\_\_\_\_\_\_\_ one electron and chlorine \_\_\_\_\_\_\_\_\_\_ one electron.

Sodium has a \_\_\_\_\_ charge and chlorine has a \_\_\_\_\_ charge. The net charge is \_\_\_\_\_\_\_\_\_\_

What is a diatomic atom? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Draw Bohr models for calcium and fluoride to demonstrate why the formula is CaF2